Cambridge Assessment International Education

Cambridge International AS & A Level

CHEMISTRY **9701/33**

Paper 3 Advanced Practical Skills 1 **February/March 2023**

MARK SCHEME

Maximum Mark: 40

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

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Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

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GENERIC MARKING PRINCIPLE 5: Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).
GENERIC MARKING PRINCIPLE 6: Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	<u>'List rule' guidance</u> For questions that require ***n*** responses (e.g. State **two** reasons …): • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked *ignore* in the mark scheme should not count towards ***n***. • Incorrect responses should not be awarded credit but will still count towards ***n***. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

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6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

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Question	Answer	Marks
1(a)	**I** **Table or list** of correct headings for 4 balance readings: • (mass of) crucible (and lid) • (mass of) crucible (and lid) and **FA 1** / contents before heating • (mass of) crucible (and lid) and residue / contents after heating (not **FA 1**) • (mass of) crucible (and lid) and residue / contents after re-heating (not **FA 1**)	**1**
	II **Recording of readings** – all 3 bullets required: • Unambiguous unit for **all** data is recorded • All weighings recorded to same number of decimal places (2 or more) • 4th weighing is ⩽ 0.05 g less than 3rd **and** is not more than 0.03 g greater than 3rd	**1**
	III **Correct subtractions of masses** Masses of **FA 1** and residue are correct.	**1**
	IV, V Accuracy (Q) marks • Examiner must check and correct (if necessary) the masses of **FA 1** used and of residue obtained by supervisor and by candidate • Calculate (mass of **FA 1**) / (mass of residue) for the supervisor to 2 dp • Write this ratio, in a ring, on the candidate script • Calculate (mass of **FA 1**) / (mass of residue) for the candidate to 2 dp • Calculate δ, the difference between these two ratios	
	Award **IV** if δ ⩽ 0.25	**1**
	Award **V** if ⩽ 0.10	**1**

Question	Answer	Marks
1(b)(i)	Correctly **calculates** of moles of zinc oxide: n(ZnO) = mass of residue / 81.4 **and** Correctly **uses** n(ZnO) to calculate moles of $ZnCO_3$•$2Zn(OH)_2$•**x**H_2O: n(ZnO) / 3 **and** both answers given to 2–4 sf	**1**
1(b)(ii)	Correctly uses **(b)(i)** to calculate M_r of $ZnCO_3$•$2Zn(OH)_2$•**x**H_2O: mass of **FA 1** / final answer **(b)(i)** **and** answer given to 2–4 sf	**1**
1(b)(iii)	Correctly calculates M_r as 324.2	**1**
1(b)(iv)	**M1** shows use of: **(b)(ii)** – **(b)(iii)** / 18	**1**
	M2 answer given correctly rounded to the nearest integer value	**1**

Question	Answer	Marks
2(a)	**I** Initial and final readings and titre recorded for rough titration **and** accurate titration details tabulated	**1**
	II Titre values recorded for accurate titrations **and** All **three** headings and units correct for accurate titrations Headings: initial / final (burette) **and** reading / volume / vol **or** reading / volume / vol at start / finish **and** titre **or** volume / **FA 3 and** added / used **and** Units: (cm^3) **or** / cm^3 **or** in cm^3 (or cm^3 by every entry)	**1**
	III All accurate burette readings are recorded to the nearest 0.05 cm^3	**1**
	IV The **final** accurate titre recorded is within 0.10 cm^3 of any other accurate titre	**1**
	V, VI, VII Award **V** if $\delta \leqslant 0.50\,cm^3$ Award **VI** if $\delta \leqslant 0.30\,cm^3$ Award **VII** if $\delta \leqslant 0.20\,cm^3$	**3**
2(b)	Candidate must average two (or more) accurate titres with total spread of not more than 0.20 cm^3. The mean value must be quoted correct to 2 dp Working must be shown or ticks must be put next to the two (or more)	**1**
2(c)(i)	**All** answers in **(c)(ii), (c)(iii), (c)(iv)** and **(c)(vi)** given to 3 or 4 sf Minimum of three answers to qualify	**1**
2(c)(ii)	Correctly calculates amount of NaOH added: (**(b)** / 1000) $\times$ 0.150	**1**

Question	Answer	Marks
2(c)(iii)	Correctly calculates amount of HCl in 1 dm^3 of **FA 2**: **(c)(ii)** × 40	**1**
2(c)(iv)	Correctly uses: **(c)(iv)** = 0.2(0) – **(c)(iii)**	**1**
2(c)(v)	Correctly balanced equation: $ZnCO_3$•2$Zn(OH)_2$•**x**H_2O + 6HCl → **3** ZnCl_2 + CO_2 + (x+**5**) H_2O	**1**
2(c)(vi)	Correctly uses: **(c)(iv)** / 6	**1**
	Correctly uses: 3.52 / (**(c)(iv)** / 6)	**1**
2(d)	*Any two for 1 mark and four for 2 marks:* • smaller titre / less **FA 3** used / smaller volume (of NaOH) used • (therefore) calculated / remaining HCl is less **or** less HCl present in (1 dm^3) **FA 2** • more HCl apparently reacted (with basic zinc carbonate) • moles of basic zinc carbonate (reacting) increases • but same mass (so M_r is smaller)	**2**

Question	Answer	Marks
3	**FA 5** is $CuCO_3$ **FA 6** is HCl	
3(a)	• Green powder • (solid) turns black • (powder) 'jumps' around the tube • Condensation / (water) droplets • Gas / CO_2 forms white ppt with limewater Any 2 bullet points = 1 mark	**2**
3(b)(i)	• effervescence **and** Mg / solid dissolves / colourless solution formed / vigorous reaction / (test-tube / mixture) gets hot	**1**
	• (gas) 'pops' with a lighted splint	**1**
3(b)(ii)	$Mg(s) + 2H^+(aq) \rightarrow Mg^{2+}(aq) + H_2(g)$	**1**
3(c)(i)	effervescence **and** blue / turquoise / green solution	**1**
3(c)(ii)	(pale) blue ppt soluble in excess giving dark / deep blue solution	**1**
3(d)(i)	silver nitrate **and** ammonia (solution)	**1**
	sodium hydroxide **and** aluminium (foil)	**1**
3(d)(ii)	**M1** minimum 2 × 2 table with unambiguous heading(s) **and** Use of both sets of reagents given in **(d)(i)**	**1**
	M2 With silver nitrate: white precipitate **and** ppt soluble / colourless solution in ammonia **and** no reaction with second set of reagents	**1**

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Question	Answer	Marks
3(e)		**2**

	cation	anion
FA 5	Cu^{2+} / copper(II)	CO_3^{2-} / carbonate
FA 6	H^+ / hydrogen	Cl^- / chloride

Any 2 correct identities = 1 mark